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**Ex Situ** Generation of Sulfuryl Fluoride for the Synthesis of Aryl Fluorosulfates

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**Supporting Information**

**ABSTRACT:** A convenient transformation of phenols into the corresponding aryl fluorosulfates is presented: the first protocol to completely circumvent direct handling of gaseous sulfuryl fluoride (SO$_2$F$_2$). The proposed method employs 1,1'-sulfonyldimidazoles as a precursor to generate near-stoichiometric amounts of SO$_2$F$_2$ gas using a two-chamber reactor. With NMR studies, it was shown that this ex situ gas evolution is extremely rapid, and a variety of phenols and hydroxylated heteroarenes were fluorosulfated in good to excellent yields.

Since 2014, aryl fluorosulfates have sparked enormous interest as they give access to a broad and powerful set of applications. This is primarily due to the specific reactivity of the fluorosulfate group.

First, the S(VI)–F bond conveys onto the sulfur center an electrophilic behavior very different from other sulfonyl halides. Owing to these unique properties, the fluorosulfate moiety is inert toward most nucleophiles but reacts cleanly with either amines (by merit of protons solvating F$^-$) or aryl silyl ethers (metathesizing into a diaryl sulfate and an extremely strong Si–F bond). This set of Lewis base mediated “click” reactions was recently disclosed by Sharpless and co-workers and baptized as SuFEx chemistry.$^1$ The SuFEx click reaction between aromatic bis(fluorosulfates) and bis(silyl ethers) can, for example, be applied for the synthesis of poly(aryl sulfates), sulfate-backboned analogs to polycarbonates which show promising mechanical properties.$^2$ SuFEx chemistry has also found intriguing applications in selective and orthogonal postpolymerization modifications$^3$ as well as in biomolecular and peptide chemistry.$^4$

Second, besides being a robust connector handle, fluorosulfates are also excellent pseudohalides in transition-metal-catalyzed cross coupling reactions. Many applications have followed, including Suzuki–Miyaura,$^5$ Negishi,$^6$ or Stille–Migita coupling,$^{67}$ alkoxyarylation,$^8$ and Buchwald–Hartwig amination.$^9$ Consequently, they are often considered as efficient triflate surrogates, albeit at a much lower production cost. Furthermore, aryl fluorosulfates are also versatile intermediates as they can be converted into aryl fluorides,$^{10}$ aryl sulfamate esters,$^{11}$ and others$^{12}$ or can even be used for the synthesis of anhydrous tetraalkylammonium fluorosulfate salts.$^{13}$

Historically, aryl fluorosulfates are mainly synthesized via four different approaches.$^{14}$ The first method relies on the pyrolysis of arenediazonium fluorosulfate salts.$^{15}$ In the other three strategies, the appropriate phenol (or phenolate) is combined with fluorosulfonic anhydride,$^{5,6,8,16}$ sulfuryl chloride fluoride,$^{17}$ or sulfuryl fluoride$^{18}$ in the presence of a base at low temperature. Unfortunately, these approaches either require highly toxic and/or expensive reagents or make use of complicated and nonreliable reaction procedures (e.g., gas condensation) which often results in low yields.

Recently, Sharpless and others described a robust, reliable, and easy-to-execute synthesis of aryl fluorosulfates.$^{12,19}$ These compounds were prepared from phenols and sulfuryl fluoride in the presence of a base, typically triethylamine. This finding has been foundational for the renewed interest in the fluorosulfate chemistry, considering that all newly discovered applications make use of this methodology.

The SO$_2$F$_2$ gas is generally introduced using a balloon, filled from a pressurized lecture bottle. Other methods rely on the use of a stock solution of sulfuryl fluoride.$^{19}$ Although sulfuryl fluoride gas is produced on an industrial scale and widely used as a fumigant, its impact on human health and the environment should not be neglected.$^{20}$ The time-weighted average exposure limit for sulfuryl fluoride has been set at 5 ppm by the Occupational Safety and Health Administration, indicating that long-term exposure is harmful for human well-being.$^{20,21}$ The gas was also recently identified as a greenhouse gas with a global warming potential of 4800 relative to carbon dioxide and an atmospheric lifetime of 36 years.$^{22}$ Moreover, the limited number of suppliers makes it challenging and expensive to obtain a gas cylinder, especially when taking into account the costs of transportation and disposal. Despite the inherent drawbacks of this reagent, all reported procedures make (in)direct use of pressurized sulfuryl fluoride gas bottles and...
thus are always associated with high cost and risks of explosion and leakage.

In summary, there is an urgent need to further develop new, convenient, and inherently safe methodologies to produce fluorosulfates, preferably starting from inexpensive and readily available commodity chemicals. Here, we report the first procedure for the on-demand ex situ generation of sulfuryl fluoride from a solid precursor for the synthesis of aryl fluorosulfates.

Our investigation began by scanning literature to identify potential sulfuryl fluoride precursors or surrogates. Inspired by the report of Prakash, sulfuryl chloride fluoride was selected as our first candidate. Therein, the SO₂ClF gas could selectively be formed from sulfuryl chloride and a fluoride source in acidic medium and subsequently be isolated after gas condensation, albeit in low yield. We assumed that translating this work to a two-chamber reactor would significantly improve this method, as the generated gas can migrate to an adjacent chamber, where it is directly consumed thus avoiding intermediate isolation. This reactor was originally designed and commercialized by the Skrydstrup group. We have recently used this setup for the ex situ generation of carbon monoxide and sulfur dioxide for applications in organic synthesis.

After an extensive optimization study (see Supporting Information (SI)), the model substrate 4-fluoro-4′-hydroxybiphenyl was successfully converted into its corresponding aryl fluorosulfate 1 and isolated in 95% yield. Unfortunately, under the optimized reaction conditions (4 equiv of SO₂Cl₂ and 6 equiv of KF in 0.6 mL of HCOOH) the transformation of the electron-rich 4-hydroxyanisole resulted in multiple chlorinated byproducts. Considering sulfuryl chloride’s known behavior as a chlorinating agent in electrophilic aromatic substitutions, we speculate that the volatile SO₂Cl₂ and/or the SO₂ClF also acted as the chlorinating agent. To further explore the scope of this methodology, a broad and diverse set of phenol derivatives was investigated (Scheme 1). First, monosubstituted electron-rich and -deficient phenols were successfully converted into their corresponding aryl fluorosulfates (3–11). These results illustrate that the fluorosulfation reaction is the rate-limiting step instead of the SO₂F₂ generation from SDI. Next, the reaction was performed at 40 °C (entries 4–5). However, when formic acid was replaced by trifluoroacetic acid, the number of equivalents of SDI and KF could be significantly reduced to 1.5 and 4.0, respectively. These conditions furnished fluorosulfate 1 in an isolated yield of 96% (entries 6–9). Again, product formation was hampered when the amount of SDI and/or KF were further decreased (entries 10–11). The optimization study was finalized by investigating the influence of the reaction duration. After 2 and 6 h, the yield was 84% and 94%, respectively (entries 12–13). We hypothesize that fluorosulfation is the rate-limiting step instead of the SO₂F₂ generation from SDI. The optimized conditions were the same as those described above.

**Table 1. Optimization of Reaction Conditions**

<table>
<thead>
<tr>
<th>Entry</th>
<th>SDI [equiv]</th>
<th>KF [equiv]</th>
<th>Acid</th>
<th>Yield [%]</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>4.0</td>
<td>6.0</td>
<td>HCOOH</td>
<td>&gt;99</td>
</tr>
<tr>
<td>2</td>
<td>3.0</td>
<td>6.0</td>
<td>HCOOH</td>
<td>&gt;99</td>
</tr>
<tr>
<td>3</td>
<td>2.0</td>
<td>6.0</td>
<td>HCOOH</td>
<td>81</td>
</tr>
<tr>
<td>4</td>
<td>2.0</td>
<td>6.0</td>
<td>HCOOH</td>
<td>82</td>
</tr>
<tr>
<td>5</td>
<td>2.0</td>
<td>8.0</td>
<td>HCOOH</td>
<td>89</td>
</tr>
<tr>
<td>6</td>
<td>2.0</td>
<td>6.0</td>
<td>TFA</td>
<td>&gt;99</td>
</tr>
<tr>
<td>7</td>
<td>2.0</td>
<td>4.0</td>
<td>TFA</td>
<td>&gt;99</td>
</tr>
<tr>
<td>8</td>
<td>1.5</td>
<td>4.5</td>
<td>TFA</td>
<td>&gt;99</td>
</tr>
<tr>
<td>9</td>
<td>1.5</td>
<td>4.0</td>
<td>TFA</td>
<td>&gt;99 (96)</td>
</tr>
<tr>
<td>10</td>
<td>1.5</td>
<td>3.0</td>
<td>TFA</td>
<td>93</td>
</tr>
<tr>
<td>11</td>
<td>1.3</td>
<td>6.0</td>
<td>TFA</td>
<td>89</td>
</tr>
<tr>
<td>12</td>
<td>1.5</td>
<td>4.0</td>
<td>TFA</td>
<td>84</td>
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<tr>
<td>13</td>
<td>1.5</td>
<td>4.0</td>
<td>TFA</td>
<td>94</td>
</tr>
</tbody>
</table>

**Reaction conditions:** Chamber A: 1,1′-sulfonyldiimidazole (SDI, 1.5 equiv), KF (4.0 equiv), and trifluoroacetic acid (TFA, 0.6 mL) at rt for 18 h. Chamber B: 4-fluoro-4′-hydroxybiphenyl (0.5 mmol), Et₃N (2.0 equiv) in dichloromethane (4 mL) at rt for 18 h. Determined by ¹⁹F NMR using trifluoroacetic acid as internal standard. Chamber A was heated to 40 °C. Isolated yield. Reaction run for 2 h. Reaction run for 8 h.

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yields (15 and 16). These slightly modified conditions were also required for the fluorosulfation of paracetamol and L-tyrosine methyl ester (17 and 18). Next, two bicyclic phenol derivatives were successfully tested (19 and 20). For the reaction of hydroquinone and phenolphthalein, 3.0 equiv of SDI and 4.0 equiv of triethylamine were added, resulting in the exclusive formation of the corresponding bis(aryl fluorosulfates) (21 and 22). The scope was finalized by the synthesis of five heteroaryl fluorosulfates (23−27). It is worth noting that, under the applied reaction conditions, anilines, aliphatic alcohols, and aliphatic amines were tolerated, as only aromatic hydroxyl groups reacted with sulfuryl fluoride (3, 16, and 18). This observation is in agreement with the results reported by the Sharpless group.1

The developed method is easily scaled up by simply using a larger two-chamber reactor. This was illustrated by synthesizing compound 11 on 5 gram scale. After an aqueous acid/base wash, the desired aryl fluorosulfate was achieved in 96% yield (see SI).

In conclusion, we have demonstrated a new, practical, and efficient way of transforming phenols into their corresponding aryl fluorosulfates in good to excellent yields. The proposed method relies on ex situ generation of sulfuryl fluoride gas from cheap and readily available commodity chemicals in a two-chamber reactor. This provides a convenient means of transforming common phenolic substances, including some drug-like and naturally occurring compounds, into reactive aromatic intermediates. Furthermore, it is easily scaled up as evidenced by the preparation of analytically pure 2-bromophenyl fluorosulfate on multigram scale using only extractive workup as the purification step. Further implementation of this promising chemistry within the larger research community can herewith be accelerated, where otherwise it might remain less explored due to the cumbersome handling of gaseous reagents. We speculate that the demand for these electrophilic functional handles will increase as they find their way into many more applications, by merit of their leaving group ability or as SuFEx partners.

ASSOCIATED CONTENT

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Supporting Information
The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.orglett.7b02522.

Experimental procedures, characterization for all compounds, decomposition study, and pressure profile (PDF)
Instructional video (AVI)

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Notes

The authors declare no competing financial interest.

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REFERENCES